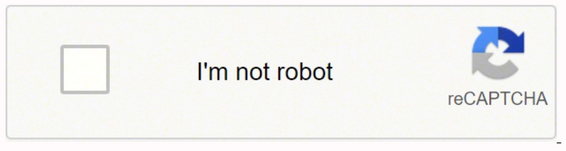

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What is the boiling point of ethanol at standard pressure

Boils aren’t usually a cause for alarm, but they can lead to a potentially worse infection, so it’s a good idea to take care of them. Boils often start out as a red bump that will continue to grow and fill with pus. Some will drain on their own. What Is a Boil?A boil will show up as a bump on the skin that fills up with pus. Sometimes they can be very painful. They can appear almost anywhere on your body, but Mayo Clinic says they typically appear in areas where hair can grow. This includes your neck, face, armpits or thighs as well as other places. Boils also typically grow in places where you sweat or can experience friction from your clothing. When a cluster of boils appears in one place it is called a carbuncle. These are typically deeper and are a more serious infection.Who Gets Boils?Some people are more prone to boils than others. This can be due to biological factors or lifestyle habits. For example, if you live with someone who has a boil or carbuncle, you could get one as well from being exposed to staph infection. People with diabetes or other various skin conditions may be prone to boils as well.What to Do for a BoilIt’s possible to treat a single boil at home on your own. Put a warm, wet cloth over the top of the boil a couple of times a day and allow it to soften. The area will probably start to drain on its own, says Healthline. Apply some light pressure without puncturing it, and eventually it should rupture and begin the draining process. Once that has happened, you can apply a dry piece of gauze to the area to keep it clean. When working with your boil, remember to keep your hands clean by frequently washing them to protect yourself from further infection.When to See a DoctorIf your boil is large, or has become a carbuncle, a physician should be seen to drain it safely and prevent the infection from spreading or getting worse. Sometimes a boil will not drain on its own, and you may want additional help to get rid of it.DiagnosisA doctor will probably be able to identify a boil just be looking at it during an office visit. He or she may decide to drain and treat it on the spot. Although it’s a rare situation, Healthline reports that the bacteria from a boil could go into the bloodstream and possibly move to other places in the body as blood poisoning. You don’t want the infection to travel, especially to your heart. Again, this is a rare situation, but it’s why some doctors will treat you immediately. MORE FROM QUESTIONSANSWERED.NET The boiling point of a substance is the temperature at which the vapor pressure of the liquid is equal to the surrounding atmospheric pressure, thus facilitating transition of the material between gaseous and liquid phases. All boiling points below are normal/atmospheric boiling points: they give the temperature at which the vapor pressure of the liquid is equal to atmospheric pressure at sea level, 1 atm. Boiling points of common materials Boiling point of water: 100 °C / 212 °F Boiling point of water (in Kelvin): 373.2 K Boiling point of ethanol: 78.37 °C / 173.1 °F Boiling point of methanol: 64.7 °C / 148.5 °F Boiling point of acetone: 56 °C / 132.8 °F Boiling point of alcohol: 78.37 °C / 173.1 °F Boiling point of nitrogen: -195.8 °C / -320.4 °F Boiling point of liquid helium: -269 °C / -452 °F Additional notes: The boiling point of salt water depends on the amount of salt added. For a 1.0 molal solution of salt (containing 58.44 grams of salt per kg of water), the boiling point is raised by 1.0 degrees Celsius. Substance BP (C) BP (F) Acetaldehyde20.8 °C69 °F Acetic Acid Anhydride139 °C282 °F Acetone50.5 °C133 °F Acetylene-84 °C-119 °F Alcohol - allyl97.2 °C207 °F Alcohol - butyl-n117 °C243 °F Alcohol - ethyl (grain, ethanol)79 °C172.4 °F Alcohol - methyl (wood, methanol)64.7 °C151 °F Alcohol - propyl97.5 °C207 °F Ammonia-35.5 °C-28.1 °F Aniline184.4 °C363 °F Benzene (Benzol)80.4 °C176 °F Butane-n-0.5 °C31.1 °F Butyric acid n162.5 °C316 °F Carbolic Acid (phenol)182.2 °C360 °F Carbon Dioxide-78.5 °C-109.3 °F Carbon Disulfide46.2 °C115 °F Carbon Tetrachloride76.7 °C170 °F Chloroform62.2 °C142 °F Decane-n173 °C343 °F Diethyl Ether34.7 °C94.4 °F Ethane-88 °C-127 °F Ether35 °C95 °F Ethyl Acetate77.2 °C171 °F Ethyl Alcohol77.85 °C172.13 °F Ethyl Bromide38.4 °C101 °F Ethylene Bromide131.7 °C269 °F Ethylene Glycol197 °C386 °F Freon refrigerant R-1123.8 °C74.9 °F Freon refrigerant R-12-29.8 °C-21.6 °F Freon refrigerant R-22-41.2 °C-42.1 °F Furfurol161.7 °C323 °F Glycerin290 °C554 °F Glycerine290 °C554 °F Heptane-n98.4 °C209.2 °F Hexane-n68.7 °C155.7 °F Jet Fuel163 °C325 °F Linseed Oil287 °C548 °F Methyl Acetate57.2 °C135 °F Methyl Iodide42.6 °C108 °F Milk100.167 °C212.3 °F Naphthalene218 °C424 °F Nitrobenzene210.9 °C412 °F Nonane-n150.7 °C302 °F Octane-n125.6 °C258 °F Olive Oil300 °C570 °F Pentane-n36 °C96.9 °F Petrol95 °C203 °F Petroleum210 °C410 °F Phenol182 °C359 °F Propane-43 °C-45 °F Propionic Acid141 °C286 °F Propylene-47.7 °C-53.9 °F Propylene Glycol187 °C368 °F Tar300 °C572 °F Toluene110.6 °C231 °F Turpentine160 °C320 °F Water (fresh)100 °C212 °F Xylene-o142.7 °C287 °F I would say that the difference in boiling points is due to an increase in external pressure. A liquid boils when its vapor pressure is equal to the external pressure - more often than not, this is the atmospheric pressure. So, at normal atmospheric pressure, or 1 atm, ethanol boils at #78.3^@"C"#. In this case, the heat you provide is enough to make the vapor pressure equal the atmospheric pressure #-># boiling takes place. If the atmospheric pressure is higher, the vapor pressure will have to be higher as well, since boiling occurs only when vapor pressure equals atmospheric pressure. As a result, you'll need to supply more heat to the sample, which means it will boil at a higher temperature. In your case, the first sample boils at #79^@"C"#, which means that the atmospheric pressure is very, very close to 1 atm, while the second sample boils at #81^@"C"#, which implies that the atmosperic pressure is a little higher than 1 atm. SIDE NOTE You can determine this mathematically by using the Clausius-Clapeyron equation #ln(P_2/P_1) = (DeltaH_"vap")/R * (1/T_1 - 1/T_2)#, where #P_1#, #P_2# - the vapor pressure at boiling point #79^@"C"# and #81^@"C"#, respectively; #DeltaH_"vap"# - the enthalpy of vaporization for ethanol - #"38.56 kJ/mol"#; #T_1#, #T_2# - the boiling points of the first and of the second sample - expressed in Kelvin. For the sake of argument, let's assume that, at 1 atm, ethanol boils at #79^@"C"#. Since #T_2# is bigger than #T_1#, you'd expect #P_2# to be bigger than #P_1#. Indeed it is #ln(P_2/1) = (38560cancel("J")/cancel("mol"))/(8.314cancel("J")/(cancel("K")cancel("mol"))) * (1/(352.15) - 1/(354.15))cancel("K")# #ln(P_2) = 0.06917 => P_2 = "1.072 atm"# Higher atmospheric pressure, higher boling point. A liquid's boiling point is the temperature at which the vapour pressure of the liquid becomes equal to the atmospheric pressure of the liquid. The liquid is transformed into a vapour at this temperature. It is at this point, with the addition of heat that the liquid transforms into vapour without any increase in the temperature. Elaborate Description of Boiling PointAs a liquid is heated, its vapour pressure increases until it is equal to the pressure of the gas surrounding it. Bubbles of vaporized liquid (gas) are then formed within the liquid, rising to the surface where they eventually burst before being released into the air. It is important to note that, at the boiling temperature, the vapour inside the bubble exerts enough pressure to keep the bubble from collapsing. The molecules of a liquid must be able to overcome the forces of attraction between them to create vapour.The liquid's boiling point depends on the surrounding pressure. It has a higher boiling point than the boiling point at normal atmospheric pressure while the fluid is at high pressure. For a given pressure, the boiling point of various liquids is different. The standard boiling point of a liquid was described in 1982 by IUPAC as the temperature at which the liquid boils under a pressure of 1 bar. For example, the boiling point of liquid a is x under normal atmospheric conditions of 1 bar but when liquid a is taken to an elevation of 10,000 ft. the boiling point of that liquid will become considerably lower than x. For any substance, the boiling point is the temperature point at which the material in the liquid stage converts into the gas phase. This occurs with water at 100 degrees centigrade. Based on the ice/water melting point and the liquid water/vapour boiling point, the Celsius scale was developed. Each material bears a boiling point of its own. The temperature starts to rise again as all the particles in the liquid phase have been converted into the gas phase, as long as heat is still being added to the surrounding system. As the temperature begins to rise, so does the kinetic energy of the particle.Factors Affecting the Boiling PointWhen it is less than 1 bar, the boiling point of a liquid is lower than its normal boiling point. When it is equal to 1 bar, the boiling point of a liquid is at its normal boiling point. When it is more than 1 bar, the boiling point of a liquid is greater than its normal boiling point. The types of molecules that a liquid is composed of determine its boiling point. If the molecules' intermolecular forces are equal to - Relatively strong, the boiling point of a liquid will be relatively high.Relatively weak, the boiling point of a liquid will be relatively low. Boiling Point of WaterIn two ways, water may boil, increase the temperature or decrease the air pressure. It is the air pressure at sea level that causes water to boil at 100°C. In a vacuum, where there is no air, water can boil at a much lower temperature. That is, body temperature would be sufficient to allow the blood to boil with water, if not for the skin that holds the blood pressurized. So, speaking of the boiling point of water, it boils at low air pressure at temperatures significantly below 100°C. In general, the melting point is defined as the point at which materials move from a solid to a liquid. The melting point of that liquid is called the temperature at which the solid changes its state to liquid at atmospheric pressure. This is the point of equilibrium at which both the liquid and solid phases occur. The substance's melting point often varies with pressure and is defined at the normal pressure.Boiling Point of EthanolEthanol has a boiling point of 78°C (173 °F), with a molecular weight (MW) of 46.Methanol Boiling PointMethanol has a Boiling point of 64.7°C (148.5 °F; 337.8 K)Boiling Point of MilkThe cow's milk's boiling point is about 203 ° F (95 ° C)Boiling Point of AlcoholThe boiling points of alcohols of equal molecular weights are much higher than those of alkanes. Ethanol, for example, has a boiling point of 78 °C (173 °F) with a molecular weight (MW) of 46, while propane (MW 44) has a boiling point of −42 °C (−44 °F).Melting PointIn general, the melting point is defined as the point at which materials move from a solid to a liquid. The melting point of that liquid is called the temperature at which the solid changes its state to liquid at atmospheric pressure. This is the point of equilibrium at which both the liquid and solid phases occur. The substance's melting point often varies with pressure and is defined at the normal pressure. The turbulent vibrations of the atoms in a solid substance overcome the forces of attraction operating within the solid at its melting point. Similar to the boiling point, the melting point of a substance is dependent upon the attractive forces taking place inside the solid. Note, the melting point of a substance in a solid state is the same as the freezing point of the same substance in a liquid state. It is at that temperature that the solid and the liquid states of the substance are in equilibrium. Melting Point of Ice32°F (0 °C) is the melting point of ice at which it, as a solid, transforms into water, a liquid.Melting Point of DiamondThe ultimate diamond melting point is around 7,280° Fahrenheit (4,027° Celsius).Did you Know?Water is unique in that it is less dense than the liquid form of the solid form, ice, which is why ice floats. Water, since it dissolves more substances than any other liquid, is considered the "universal solvent" This implies that it brings along valuable chemicals, minerals, and nutrients everywhere water goes, either through the earth or through our bodies. The neutral pH of pure water is 7, which is neither acidic (less than 7) nor basic (greater than 7). It is very sticky, meaning water molecules bind to each other, the water molecule is extremely cohesive. Of the non-metallic liquids, water is the most coherent. The water molecule is highly adhesive.Learn more about the concept of the natural phenomenon of a boiling point along with examples at Vedantu. Get excellent study material to understand this phenomenon and understand it well. Make your conceptual foundation stronger and answer questions easily.

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